Writing formulas and naming acids

Recognizing acids

Acids are a special group of ionic compounds where hydrogen acts like a metal. They always follow the form "**HX**," where H is always hydrogen and X is always some nonmetal or polyatomic ion (a "red" element).

Since *all* acids have hydrogen as the first element, it does not need to be used in the name. So, acids are named based on their anions, or nonmetal elements.

The *X* controls the name.

Here are some examples:

Formula	Category 1 name	Acid name	Rule
HCl	Hydrogen chlor***ide***	***Hydro***chlor***ic*** acid	"-***ide***" endings turn into "***hydro-ic***" names
H_2S	Hydrogen sulf***ide***	***Hydro***sulfur***ic*** acid	"-***ide***" endings turn into "***hydro-ic***" names
HNO_3	Hydrogen nitr***ate***	Nitr***ic*** acid	"-***ate***" endings turn into "-***ic***" names
H_3PO_4	Hydrogen phosph***ate***	Phosphor***ic*** acid	"-***ate***" endings turn into "-***ic***" names
H_2SO_2	Hydrogen sulf***ite***	Sulfur***ous*** acid	"-***ite***" endings turn in "-***ous***" names
HClO	Hydrogen hypochlor***ite***	Hypochlor***ous*** acid	"-***ite***" endings turn in "-***ous***" names

If the name of the X element ends with "-***ate***," the acid name ends in "-***ic***."
If the name of the X element ends with "-***ite***," the acid name ends in "-***ous***."
If the name of the X element ends with "-***ide***," the acid name use "***hydro-ic***."

Just remember: -ate to -ic
-ite to -ous
-ide to hydro-ic